

		Marks
Question 29 — Shipwrecks and Salvage (25 marks)		
(a)	(i) Identify the main metal used to construct ships.	1
	(ii) Although aluminium is a very reactive metal, with a very low reduction potential, it is used in many structures exposed to oxidising conditions. Explain why aluminium can be used in this way.	2
(b)	(i) Give an example of a metal commonly used as a sacrificial anode.	1
	(ii) Explain why sacrificial anodes are added to metal-hulled ships.	3
(c)	Describe the effect of adding other elements to iron on the properties and uses of steels.	5
(d)	(i) Define <i>corrosion</i> .	1
	(ii) Outline a procedure that could be used to compare the corrosion rates of different metals or alloys in the school laboratory.	2
	(iii) Describe ways in which accuracy and reliability could be improved in the procedure described in part (ii).	3
(e)	For ONE specific metal, evaluate the steps that can be used to clean, stabilise and preserve artefacts recovered from shipwrecks.	7