

Chemistry

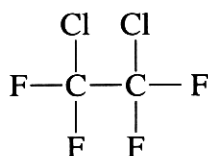
Section I – Part B (continued)

Marks

Question 25 (6 marks)

- (a) What is the systematic name of the CFC in the diagram?

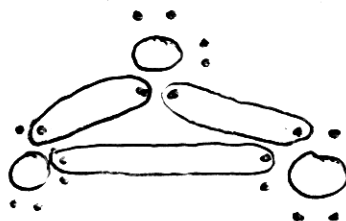
1



dichlorotetrafluorocarbon

- (b) Identify the bonding within ozone, using a Lewis electron-dot diagram.

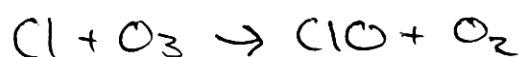
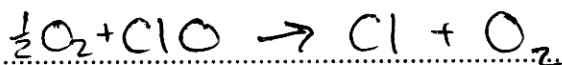
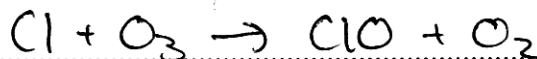
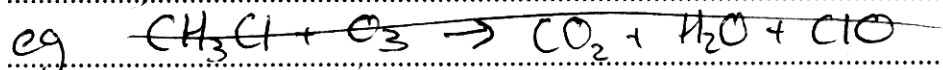
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- (c) Discuss how CFCs damage the ozone layer, using relevant equations.

3

CFC continue to destroy the ozone layer because the reactant, Cl, is never used up, causing a chain reaction.



\therefore chain reaction, reactant never used up.