ARD OF STUDIES QUESTION 29 - Shipwrecks and Salvage (a) Galvanic Cell $Fe^{2+} + 2e^- \rightleftharpoons Fe(s) = -0.44V$ (ii) $(u^{1+} + 2e^{-} \rightleftharpoons (u^{(5)}) = 0.34V$ b) Bronsten and Lowry both increased our understanding of electron transfer reactions. (c) (i) (ii)d) (i) . We placed a nails in a test-tubes. water (aistilled) corks · Added acids, Marson, jovered some with approximity and left some opened. Boiled one with water to see affect of heat. · we left them for a few days to see the rate of corrosion. W 5 boiled water then corked acid water (ii) For something to corrode, it needs O2 to be present. With the acid we used HCI, O2 is not present therefore corroded less than the one with the H2O. (e) It is colder at the bottom of the ocean. less of There is less oxyget also.

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ARD OF STUDIES B (But still a metal will corrode. . .

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