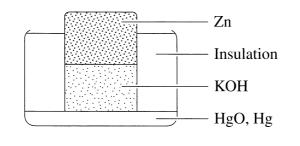
Question 27

2010 HSC Chemistry

Question 27 (2 marks)

The diagram shows a particular cell with relevant half equations.



$\operatorname{Zn}(s) + 2\operatorname{OH}^{-}(aq)$	\rightarrow	ZnO	(s) +	$H_2O(l)$	+ 2e ⁻
$HgO(s) + H_2O(l) +$	2e ⁻	\rightarrow	Hg(<i>l</i>)) + 20	$H^{-}(aq)$

Identify the anode, cathode and electrolyte for this cell.

electrolyte = KOH	Zince anode = zinc
elletrolyte - KOH	(athode = Hg 0)
$\gamma \gamma $	electrolyte = KOH
0-	0-

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