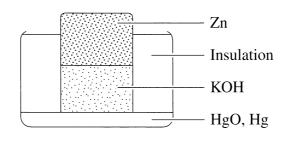
## **Question 27**

## 2010 HSC Chemistry

## Question 27 (2 marks)

The diagram shows a particular cell with relevant half equations.



$Zn(s) + 2OH^{-}(aq)$	$\rightarrow$	ZnO	(s) + H	$_{2}O(l) + 2e^{-}$
$HgO(s) + H_2O(l) +$	2e <sup>-</sup>	$\rightarrow$	Hg(l)	+ 2OH <sup>-</sup> ( <i>aq</i> )

Identify the anode, cathode and electrolyte for this cell.

Zn is oxidised at chartle	
Had is reduced at esthude	
electrolyte is KOH	
/	